

108 學年度學生轉系考試試題參考答案疑義釋疑

考科	題號	釋疑答覆	釋疑結果
英文	7	經過確認後發現 aphasia 及 dyslexia 兩字皆 unable or difficult to read and write 的含意，所以答案 A 及 C 都算正確答案	A 或 C
	29	According to the description that “Edd Hammill found that mosquito population boom because of mosquitoes predators have not kept pace with the evolution while conducting research in orange plantations in northern Costa Rica. “, it is concluded that “A farm in northern Costa Rica” would be the place where more mosquito can be found.	維持原答案 B
	35	本文章一開始提及 body shop 的發源，並在中間介紹了它於全盛期所發生及所支持的一些社會活動。而非針對 Anita Roddick 個人生平的敘述，所以答案 A 才是最適合的答案	維持原答案 A

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普通 生物 學	7	According to Campbell biology -concepts and connections 8 th ed. The maximum ATP after “oxidative” phosphorylation is 32 ATP (page 138, Figure 6.12). The difference between this figure and the sentence that the examinee provided is due to different authors use the term in flexible context (apparently, Campbell biology-concepts and connections used oxidative phosphorylation to represent this whole step). However, despite the terms used by different authors, from the answers, the answer A) is clearly the best answer of the all.	維持原答案
	40	Fat globules can be digested by bile salts into fat droplets. Droplets then are digested into fatty acids glycerol. The examinee provided the digestion pathway for fat in general, thus is irreverent to this question.	維持原答案

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考科	題號	釋疑答覆	釋疑結果
普通 化學	5	$\text{Ag}^+ + \text{S}_2\text{O}_3^{2-} \rightleftharpoons \text{Ag}(\text{S}_2\text{O}_3)^- \quad K_1 = 7.4 \times 10^8$ $\text{Ag}(\text{S}_2\text{O}_3)^- + \text{S}_2\text{O}_3^{2-} \rightleftharpoons \text{Ag}(\text{S}_2\text{O}_3)_2^{3-} \quad K_2 = 3.9 \times 10^4$ $[\text{Ag}^+]_0 = \frac{(150\text{ml})(1.0 \times 10^{-3}\text{M})}{(150+200)\text{ml}} = 4.29 \times 10^{-4}\text{M}$ $[\text{S}_2\text{O}_3^{2-}]_0 = \frac{(200\text{ml})(5.0\text{M})}{(150+200)\text{ml}} = 2.86\text{M}$ <p>Net reaction : $\text{Ag}^+ + 2\text{S}_2\text{O}_3^{2-} \rightarrow \text{Ag}(\text{S}_2\text{O}_3)_2^{3-}$</p> <p>Before reaction: $4.29 \times 10^{-4}\text{M} \quad 2.86\text{M} \quad 0$</p> <p>After reaction: $\sim 0 \quad 2.86 - 2(4.29 \times 10^{-4}) \quad 4.29 \times 10^{-4}\text{M}$</p> <p style="text-align: center;">$\approx 2.86\text{M}$</p> <p>We can calculate the concentration of $\text{Ag}(\text{S}_2\text{O}_3)^-$ from K_2 :</p> $3.9 \times 10^4 = K_2 = \frac{[\text{Ag}(\text{S}_2\text{O}_3)_2^{3-}]}{[\text{Ag}(\text{S}_2\text{O}_3)^-][\text{S}_2\text{O}_3^{2-}]} = \frac{4.29 \times 10^{-4}}{[\text{Ag}(\text{S}_2\text{O}_3)^-]} \quad (2.86)$ $[\text{Ag}(\text{S}_2\text{O}_3)^-] = 3.8 \times 10^{-9}\text{M}$ <p>We can calculate $[\text{Ag}^+]$ from K_1:</p> $7.4 \times 10^8 = K_1 = \frac{[\text{Ag}(\text{S}_2\text{O}_3)^-]}{[\text{Ag}^+][\text{S}_2\text{O}_3^{2-}]} = \frac{3.8 \times 10^{-9}}{[\text{Ag}^+]} \quad (2.86)$ $[\text{Ag}^+] = 1.8 \times 10^{-18}\text{M}$	維持原答案